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Evaluation of the thermochemical properties of Ho–Cd intermetallic compounds using electrochemical techniques

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The electrochemical formation of HoCd₆, HoCd_{45/11}, HoCd₃, HoCd₂ and HoCd, on a Cd coated tungsten electrode in the eutectic LiCl–KCl has been evaluated in a qualitative and quantitative manner using electrochemical techniques (*i.e.* cyclic voltammetry, anodic stripping voltammetry and open circuit chronopotentiometry) over the temperature range of 673–823 K. The formation energy of each intermetallic compound, as well as the global formation constants were calculated for the first time. The activities and relative partial molar Gibbs energies of Ho and Cd were also obtained for HoCd₆, HoCd_{45/11}, HoCd₃, HoCd₂ and HoCd. At the Cd distillation temperature, the molar Gibbs free energy of the intermetallic compounds achieves the minimum value for HoCd₂ and HoCd, suggesting that the dissolution of U or Pu from the U–Pu–Ho–Cd alloy into HoCd₂ and HoCd should be considered.

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Introduction

Pyro reprocessing fuel cycle technology appears as an encouraging candidate to carry out an advanced fuel cycle with reduced loads of long life radioactive waste and no proliferation risk. In a closed fuel cycle scenario, the radiotoxicity of the spent nuclear fuel can be reduced separating the Long-Lived Fission Products (LLFP) and Minor Actinides (MAs), which are burned and converted into short-lived isotopes. This strategy, known as partitioning and transmutation (P&T), could be carried out in dedicated accelerator driven systems or integrated in the advance fuel cycle of a fast reactor system.¹

In pyrochemical reprocessing, the separation takes place in a molten salt media (*i.e.* chlorides or fluorides) due to their advantageous properties, mainly their high radiation stability and thermal resistance, which allows the processing of spent fuels with short cooling times, and not being neutron moderators, a large amount of fissile material can be handled.² In pyrochemical reprocessing, the proposed separation strategies are: (i) electrodeposition on inert (generally W or Mo) and reactive (*e.g.* Cd, Bi, Al) cathodes, and (ii) molten salt reductive extraction using a liquid metal. Special attention must be given to lanthanides, Lns, the most awkward FPs to be separated from the Ans, since their large capture cross sections would spoil the neutron economy of the reactor core in the transmutation stage.³ It merits to be mentioned, that during the separation stage the Lns accumulate in the molten salt media modifying its characteristics and contaminating

the final product. Therefore, in order to ensure a longer lifetime of the media, it is required to remove the Lns from the salts after the selective extraction of Ans,⁴ activities in which our team has been engaged since 2000.

Cd is assumed to be a typical liquid metal to be used in pyrochemical reprocessing, since Cd has a low melting point (594 K) and the activities of Ans and Lns in the metallic phase are very small.⁵ In this way, the electrochemical behaviour of some Ans and Lns has been examined using a liquid cadmium cathode.^{5–19} Nevertheless, there are still many of them that need to be explored in terms of electrochemical and thermodynamic properties. This is the case of holmium, a heavy lanthanide object of the present paper, in which its electrode reactions at a Cd coated W electrode, were investigated, with the aim to calculate the Gibbs free energies of formation of the different Ho–Cd intermetallic compounds, as well as the activities and relative partial molar Gibbs energies of Ho and Cd, by means of electrochemical techniques. The paper describes the methods used and the obtained results.

Experimental

Chemical storage, electrolytic bath and sample preparation were handled in a MBraun glove box in purified Ar atmosphere. The electrolytic bath, 80 g of the eutectic LiCl–KCl, was introduced in a high purity 100 cm³ alumina crucible placed in a cylindrical quartz cell and then fused under vacuum and purified with HCl(g) following the methodology described elsewhere.^{14,17}

Holmium and cadmium ions were introduced into the bath in the form of anhydrous HoCl₃ and CdCl₂ (Sigma Aldrich 99.99%). The experiments were performed under a dry Ar

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atmosphere, and the temperatures were controlled with a nickel–chromium thermocouple.

Different working electrodes have been used: (i) a tungsten wire of 1.0 mm diameter, (ii) a liquid Cd bulk electrode, prepared by placing some granules of Cd (99.999%; Sigma Aldrich) in a J shaped Pyrex tube, establishing the contact by means of a tungsten wire immersed in the Cd phase through the Pyrex tube, and (iii) a tungsten wire of 1.0 mm diameter (Sigma Aldrich 99.99%) coated with Cd(0). As counter electrodes, a glassy carbon rod (Sofagel) of 3 mm diameter and a 1.5 mm diameter tungsten wire (Alfa Aesar 99.95%) have been used. All the potentials were referred to the equilibrium potential adopted by a silver wire dipped into an AgCl solution (0.75 mol kg⁻¹) in the LiCl–KCl mixture, contained in a Pyrex tube.

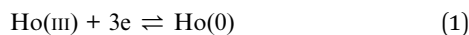
The electrochemical measurements were performed with a potentiostat/galvanostat, either a PAR EG&G Model 273A controlled by the EG&G M270 software or an Autolab PGSTAT 320N controlled by the NOVA 1.10 research software.

Results and discussion

Electrode reactions of Ho(III)/Ho system in the reactive Cd pool electrode

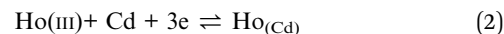
Fig. 1 displays the electrochemical spectra obtained with a solution of dissolved HoCl₃ in the eutectic bath, probably as the entity HoCl₆³⁻ and hereafter Ho(III). Curve (1) has been registered at the liquid Cd pool electrode, whereas curve (2) at the W one, in order to compare with the liquid substrate because no alloys exist for the W–Ho binary systems. Another electrochemical spectra, the red dotted curve (3), has also been recorded at the Cd pool electrode in the absence of Ho(III), being the cathodic limit attributed to the Li deposition in Cd,⁵ and the anodic one to the anodic dissolution of the own electrode.

If the substrate is an inert material, as we have reported previously,²⁰ the reaction proceeds as a simple deposition/dissolution mechanism:



whereas, at the Cd pool the mechanism seems to be different. The voltammogram consist in a cathodic wave, A^c, related to

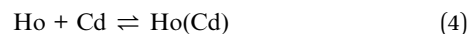
a best defined anodic one, A^a, suggesting by their shape that they are due to the deposition of Ho in Cd, Ho(Cd), and the oxidation of the Ho dissolved in Cd, respectively. A similar behavior to those reported for other Lns and Ans in liquid cathodes.^{5,11,17,22–24}



From a thermodynamic point of view, the potential shift of the electrochemical system in the anodic direction *versus* the W electrode, is caused by the decrease in the activity of Ho, which is dissolved in the Cd phase, as described in eqn (3)

$$E_{\text{eq}} = E_{\text{Ho(III)/Ho}}^0 + \frac{RT}{3F} \ln \frac{a_{\text{Ho(III)in solution}}}{a_{\text{Ho(Cd)}}} \quad (3)$$

On the other hand, as in the case of the Hg electrode in aqueous solutions,²² the reaction scheme can also be viewed, as an electrochemical–chemical, EC, process, described by eqn (1) and (4)



The chemical step (4) can be considered as an equilibrium step, being the product of the charge transfer (1) stabilized by this step, and hence the $E_{1/2}$ is shifted in the positive direction.

It merits to be mentioned that if the amount of electrodeposited Ho surpasses its solubility in Cd, the electrode homogeneity is lost, appearing separated solid phases (*i.e.* intermetallic compounds), a phenomenon that can be observed in Fig. 2. When the inversion potential achieves cathodic values greatest than –1.85 V, in addition to the cathodic wave A^c, an important cathodic current is noticeable, appearing in the reverse scan a new signal, labelled B^a, at around –1.72 V, with the shape of a dissolution peak. Therefore, part of the cathodic current would be caused by the deposition of an intermetallic compound, probably HoCd₆, which according to the reported data²¹ is the intermetallic with the richer content in Cd, being the peak B^a produced by its reoxidation. Hence, the transformations should be as follows:

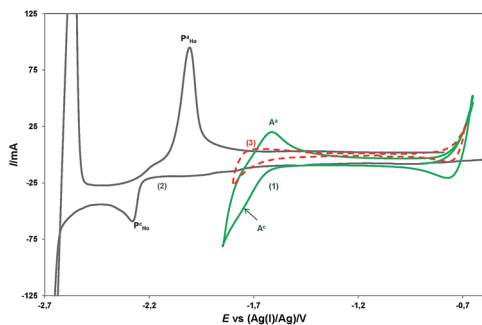


Fig. 1 CVs registered with a HoCl₃ solution ($C_0 = 1.17 \times 10^{-4}$ mol cm⁻³) at a Cd pool, curve (1), and a W electrode, curve (2). Curve (3) shows the CV obtained at the Cd pool electrode in absence of HoCl₃. Scan rate: 100 mV s⁻¹; temperature 698 K.

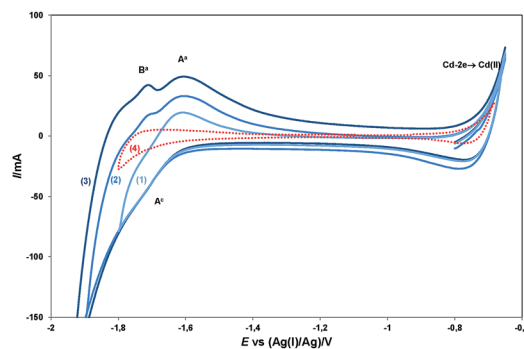
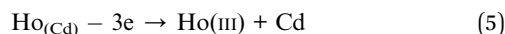
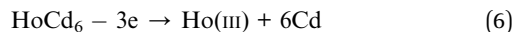


Fig. 2 CVs obtained with a Ho(III) solution ($C_0 = 1.17 \times 10^{-4}$ mol cm⁻³) at a Cd pool electrode (scan rate: 100 mV s⁻¹, T : 698 K) at different inversion potentials, curves (1–3). Curve (4) cyclic voltammogram registered with the Cd pool electrode in absence of HoCl₃.

Peak A^a:Peak B^a:

Electrode reaction of Ho(III) at cadmium film electrodes

According to the literature²¹ holmium can form five intermetallic compounds with Cd (*i.e.* HoCd₆, Ho₁₁Cd₄₅, HoCd₃, HoCd₂ and HoCd).

With the aim to obtain thermodynamic information concerning their formation, the electrochemical behaviour of Ho(III) on a tungsten electrode coated with a thin film of Cd, CdFE, has been studied by using electrochemical techniques. Due to its small volume, the CdFE, purposefully made to be thin, has been chosen instead of the Cd pool. Since the thickness of the film is smaller than the diffusion layer ones, the semi-infinite diffusion model is not valid in the CdFE, and distinct non-homogeneous phases will be formed during deposition.²²

Detection of intermetallic compounds

Fig. 3 shows the electrochemical spectrum (blue line) obtained with a Cd(II) solution on a W electrode at 698 K, which has been divided in two consecutive zones. In zone (I) a group of signals (peaks P_{Cd}^c/P_{Cd}^a) is observed, which corresponds to the Cd(II) reduction to Cd(0) following a single-step process exchanging 2 electrons.^{17,25} In zone II, and before the massive reduction of lithium, 3 groups of signals I_c/I_a, II_c/II_a, and III_c/III_a are evident, that can be attributed to the formation of Li–Cd and/or K–Cd alloys by underpotential deposition (UPD) of the alkaline on the liquid cadmium already deposited and coating the tungsten electrode. In order to facilitate a comparison, in Fig. 3 it has also been represented the spectrum recorded on a tidy W electrode with a Ho(III) solution at the same temperature (red line)

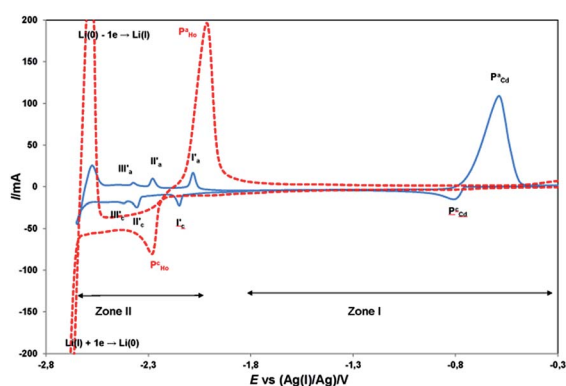


Fig. 3 CVs registered at 200 mV s⁻¹ and 673 K on a W electrode with a Cd(II) solution ($C_0 = 7.85 \times 10^{-5} \text{ mol cm}^{-3}$) (continuous blue line). Cyclic voltammograms for the reduction of a HoCl₃ solution in the eutectic LiCl–KCl ($C_0 = 1.17 \times 10^{-4} \text{ mol cm}^{-3}$) (dotted red line).

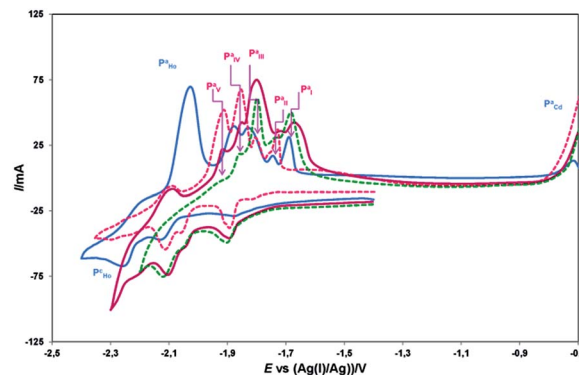


Fig. 4 CVs registered at different inversion potentials at 100 mV s⁻¹ and 723 K with a solution containing Ho(III) $1.14 \times 10^{-4} \text{ mol cm}^{-3}$ on a CdFE electrogenerated “*in situ*” at -1.4 V ($[\text{CdCl}_2] = 1.16 \times 10^{-5} \text{ mol cm}^{-3}$).

exhibiting the aforementioned electrochemical signals (P_{Ho}^c/P_{Ho}^a). Taken into account all this information, the formation of a Cd film can be carried out at the surface of a tidy W electrode, “*ex situ*” (*i.e.* in a LiCl–KCl bath containing CdCl₂) or “*in situ*” (*i.e.* in a LiCl–KCl cell containing both Ho(III) and Cd(II)), by potentiostatic electrolysis at more cathodic potentials than -0.82 V (*e.g.* -1.4 V) for a little period.

Fig. 4 shows typical examples of the cyclic voltammograms registered with a CdFE electrogenerated “*in situ*” at 698 K. A series of new redox couples can be observed between the electrochemical signals P_{Ho}^c/P_{Ho}^a and P_{Cd}^a, which are attributed to the formation/dissolution of different Ho–Cd intermetallic compounds. The anodic peaks, P_{IV}^a, P_{III}^a, P_{II}^a, and P_I^a, better defined than their corresponding cathodic ones, are due to the dissolution of Ho as Ho(III) into the LiCl–KCl, with the corresponding transformation of one intermetallic compound into another one with less Ho amount.

Anodic stripping voltammetry, ASV, is a simple and easy “*in situ*” method of redissolution that has been found to be particularly useful for phase identification in metallic alloys.

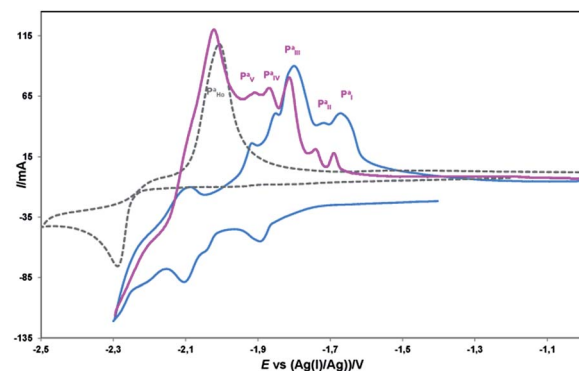


Fig. 5 Pink line, ASV obtained at 723 K with a solution of Ho(III) $1.14 \times 10^{-4} \text{ mol cm}^{-3}$ on a Cd coated W electrode, ($E_d = -2.3 \text{ V}$, $t_d = 20 \text{ s}$, scan rate 50 mV s^{-1}). Blue line, CV registered with the same solution at the same temperature on a Cd coated W electrode scan rate 100 mV s^{-1} . Grey line, CV obtained on a tidy W electrode.

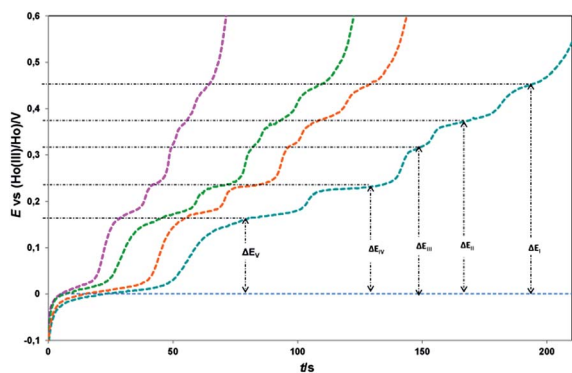


Fig. 6 Examples of the open circuit transient curves (OCP) for a W based Cd film electrode, electrogenerated *in situ*, after depositing Ho at -2.25 V vs. (Ag(l)/Ag) for 20 (–), 40 (–), 60 (–) and 80 s (–) at 723 K.

The technique is very sensitive to the interaction between the components, revealing the presence of different phases.²⁶ Dissolution of Ho–Cd intermetallic compounds displays an anodic current pattern with a series of dissolution peaks, reflecting successive dissolution of Ho, the less noble component, from each phase separately. This behaviour can be explained by the energies of formation of the intermetallic compounds, rendering different dissolution potentials.

To obtain qualitative information, samples of thin layers of Ho–Cd alloys were prepared as follows: The CdFE was electrogenerated by cathodic deposition at -1.4 V, thereafter the potential was maintained at a value corresponding to the deposition of Ho at a tidy W electrode for synthesizing HoCd, the intermetallic compound with the richer content in Ho. Following the deposition time, the potential was scanned anodically. Fig. 5 illustrates an example of the stripping voltammograms. The ASV (pink line) shows, in addition to the

Table 1 Thermodynamic properties of Ho for Ho–Cd intermetallic compounds in two-phase coexisting states at various temperatures

T (K)	Number of data	$\Delta E_{\text{Ho}(m)/\text{Ho}}$ (V)	$\Delta \tilde{G}_{\text{Ho}}$ (kJ (mol Ho) ⁻¹)	a_{Ho}
In the two-phases coexisting state between HoCd and HoCd₂				
673	5	0.167 ± 0.007	-48.4 ± 2.1	$(1.9 \pm 0.7) \times 10^{-4}$
698	12	0.164 ± 0.005	-47.4 ± 1.4	$(3.0 \pm 0.7) \times 10^{-4}$
723	23	0.161 ± 0.005	-46.6 ± 1.6	$(4.5 \pm 1.2) \times 10^{-4}$
748	19	0.169 ± 0.009	-49.1 ± 2.7	$(4.2 \pm 2.1) \times 10^{-4}$
773	65	0.150 ± 0.009	-43.5 ± 2.6	$(1.2 \pm 0.5) \times 10^{-3}$
823	11	0.152 ± 0.005	-44.0 ± 1.4	$(1.7 \pm 0.3) \times 10^{-3}$
In the two-phases coexisting state between HoCd₂ and HoCd₃				
673	10	0.240 ± 0.014	-69.5 ± 4.1	$(5.2 \pm 4.3) \times 10^{-6}$
698	11	0.222 ± 0.006	-64.1 ± 1.6	$(1.7 \pm 0.5) \times 10^{-5}$
723	48	0.216 ± 0.008	-62.4 ± 2.3	$(3.3 \pm 1.3) \times 10^{-5}$
748	45	0.228 ± 0.009	-66.0 ± 2.5	$(2.7 \pm 1.0) \times 10^{-5}$
773	65	0.212 ± 0.009	-61.3 ± 2.6	$(7.8 \pm 3.0) \times 10^{-5}$
823	23	0.227 ± 0.009	-65.7 ± 2.6	$(7.3 \pm 3.1) \times 10^{-5}$
In the two-phases coexisting state between HoCd₃ and Ho₁₁Cd₄₅				
673	11	0.307 ± 0.004	-88.9 ± 1.1	$(1.3 \pm 0.3) \times 10^{-7}$
698	13	0.309 ± 0.005	-89.5 ± 1.5	$(2.1 \pm 0.5) \times 10^{-7}$
723	50	0.301 ± 0.009	-87.1 ± 2.5	$(5.6 \pm 2.3) \times 10^{-7}$
748	38	0.312 ± 0.009	-90.2 ± 2.6	$(5.5 \pm 2.0) \times 10^{-7}$
773	71	0.298 ± 0.013	-86.3 ± 3.7	$(1.7 \pm 0.9) \times 10^{-6}$
823	18	0.300 ± 0.009	-86.9 ± 2.6	$(3.3 \pm 1.3) \times 10^{-6}$
In the two-phases coexisting state between Ho₁₁Cd₄₅ and HoCd₆				
673	12	0.376 ± 0.004	-108.8 ± 1.3	$(3.8 \pm 1.0) \times 10^{-9}$
698	15	0.361 ± 0.011	-104.4 ± 3.2	$(1.7 \pm 0.7) \times 10^{-8}$
723	52	0.356 ± 0.006	-102.9 ± 1.8	$(3.9 \pm 1.1) \times 10^{-8}$
748	33	0.359 ± 0.008	-104.0 ± 2.4	$(6.0 \pm 2.9) \times 10^{-8}$
773	82	0.346 ± 0.009	-100.0 ± 2.5	$(1.9 \pm 0.7) \times 10^{-7}$
823	22	0.351 ± 0.010	-101.6 ± 3.0	$(3.9 \pm 1.8) \times 10^{-7}$
In the two-phases coexisting state between HoCd₆ and Cd				
673	7	0.463 ± 0.006	-134.0 ± 1.7	$(4.2 \pm 1.4) \times 10^{-11}$
698	11	0.436 ± 0.009	-126.2 ± 2.5	$(3.9 \pm 1.1) \times 10^{-10}$
723	25	0.441 ± 0.010	-127.7 ± 2.9	$(6.9 \pm 3.6) \times 10^{-10}$
748	23	0.448 ± 0.013	-129.6 ± 3.7	$(1.0 \pm 0.5) \times 10^{-9}$
773	29	0.429 ± 0.018	-124.1 ± 5.2	$(5.5 \pm 3.9) \times 10^{-9}$
823	3	0.438 ± 0.002	-126.7 ± 0.6	$(9.2 \pm 0.8) \times 10^{-9}$

peaks due to Ho and Cd, five dissolution peaks labelled P_V^a , P_{IV}^a , P_{III}^a , P_{II}^a and P_I^a , in agreement with the anodic peaks obtained by cyclic voltammetry.

Open-circuit chronopotentiometry has proven to be a convenient technique for the study of potential alloy formation in both qualitative and quantitative mode.^{12-17,27-29} In the present study, the measurements were conducted in the following manner, once the CdFE was electrogenerated, the intermetallic compounds were electroformed by applying more cathodic potentials. Then, the open circuit potential was measured disconnecting the potentiostatic control. During this process, a potential plateau is observed when the electrode is composed of a mixture of alloys of definite composition (*i.e.* a two phase coexisting state).

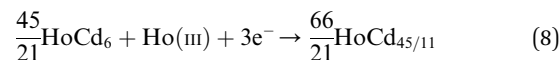
The equilibrium potentials were converted to electromotive forces (emf) against Ho(0), making them possible to estimate the relative molar Gibbs energies of Ho in the intermetallics and their Gibbs energies of formation following the methodology previously developed.^{12,14,27,29} Fig. 6 displays examples of the open circuit potential transient curves at 723 K, being possible to observe five potential plateaux labelled: I, II, III, IV and V, in agreement with the formation of five intermetallic compounds, and ratifying the results obtained by ASV and cyclic voltammetry.

Accordingly, the plateaux can be attributed to the following reactions:

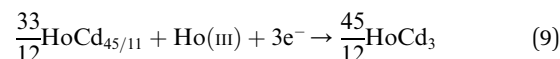
Plateau I:



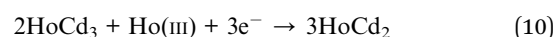
Plateau II:



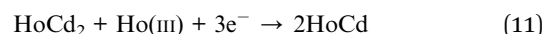
Plateau III:



Plateau IV:



Plateau V:



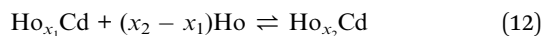
In order to confirm the reproducibility of the experiment, the measurements were repeated several times. The uncertainty was given by standard deviation for each measurement.

Table 2 Standard Gibbs free energies of Ho–Cd intermetallic compounds at various temperatures

Formation reaction	Equation for ΔG_f^0	T (K)	ΔG_f^0 (kJ mol ⁻¹)
Ho + 6Cd → HoCd ₆	$-3F\Delta E_I^*$	673	-134.0 ± 1.7
		698	-126.2 ± 2.5
		723	-127.7 ± 2.9
		748	-129.6 ± 3.7
		773	-124.1 ± 5.2
		823	-126.7 ± 0.6
Ho + $\frac{45}{11}$ Cd → HoCd _{$\frac{45}{11}$}	$\frac{21}{66} \left(\frac{45}{21} \Delta G_{f, \text{HoCd}_6}^0 - 3F\Delta E_{II} \right)$	673	-126.0 ± 0.4
		698	-119.3 ± 1.0
		723	-119.8 ± 0.6
		748	-121.4 ± 0.8
		773	-116.4 ± 0.8
		823	-118.7 ± 0.9
Ho + 3Cd → HoCd ₃	$\frac{12}{45} \left(\frac{33}{12} \Delta G_{f, \text{HoCd}_{\frac{45}{11}}}^0 - 3F\Delta E_{III} \right)$	673	-116.1 ± 0.3
		698	-111.3 ± 0.4
		723	-111.1 ± 0.7
		748	-113.1 ± 0.7
		773	-108.4 ± 1.0
		823	-110.2 ± 0.7
Ho + 2Cd → HoCd ₂	$\frac{1}{3} (2\Delta G_{f, \text{HoCd}_3}^0 - 3F\Delta E_{IV})$	673	-100.6 ± 1.4
		698	-95.6 ± 0.5
		723	-94.9 ± 0.8
		748	-97.4 ± 0.8
		773	-92.7 ± 0.9
		823	-95.4 ± 0.9
Ho + Cd → HoCd	$\frac{1}{2} (\Delta G_{f, \text{HoCd}_2}^0 - 3F\Delta E_V)$	673	-74.5 ± 1.1
		698	-71.5 ± 0.7
		723	-70.7 ± 0.8
		748	-73.2 ± 1.3
		773	-68.1 ± 1.3
		823	-69.7 ± 0.7

Thermodynamic properties of Ho–Cd intermetallic compounds

If a mixture of intermetallic compounds (e.g. Ho_{x₁}Cd and Ho_{x₂}Cd) constitutes the electrode surface, the activity of Ho is fixed by the equilibrium:



and a constant value of the emf is observed along the complete transformation of the definite Ho_{x₁}Cd into the definite Ho_{x₂}Cd.

However, if the electrode surface acquires the exact composition of an intermetallic, Ho_{x₂}Cd for example, then the emf varies, from the value of the plateau corresponding to the mixture of Ho_{x₁}Cd and Ho_{x₂}Cd to the value of the plateau of Ho_{x₂}Cd and Ho_{x₃}Cd mixture.

The relative partial molar Gibbs free energies, and activities of Ho, $\Delta\tilde{G}_{\text{Ho}}$ and a_{Ho} , in the Ho–Cd intermetallic compounds, where calculated from the obtained emf, ΔE vs. (Ho(III)/Ho)/V, through eqn (13) and (14):

$$\Delta\tilde{G}_{\text{Ho}} = -3FE \quad (13)$$

$$a_{\text{Ho}} = \exp\left(\frac{-3FE}{RT}\right) \quad (14)$$

The estimated thermodynamic data are summarized in Table 1.

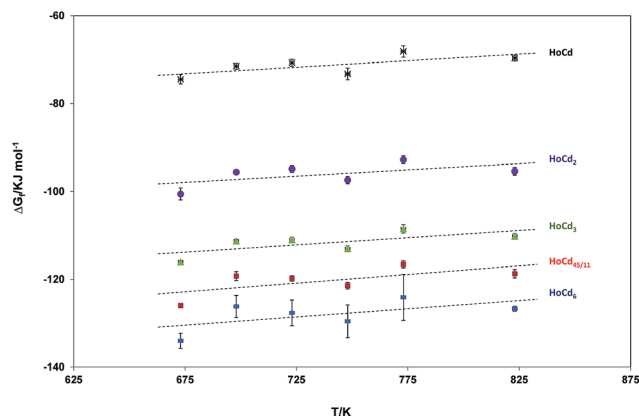


Fig. 7 Variation of the ΔG_f^0 of the different HoCd_x intermetallic compounds: (*) HoCd, (●) HoCd₂, (▲) HoCd₃, (■) HoCd_{45/11}, (◆) HoCd₆.

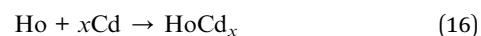
On the other hand, the standard Gibbs energies of formation ΔG_f^0 of an alloy of composition Ho_{x₂}Cd is related to that of a Ho_{x₁}Cd alloy by the relation:

$$\Delta G_f^0(\text{Ho}_{x_2}\text{Cd}) = -3F \int_{x_1}^{x_2} \Delta E(x) dx + \Delta G_f^0(\text{Ho}_{x_1}\text{Cd}) \quad (15)$$

being x_1 and x_2 the stoichiometric coefficients of Ho (i.e. $x_1, x_2 = 1, 1/2, 1/3, 11/45$ and $1/6$, with $x_1 > x_2$).

The final derived equations for ΔG_f^0 , and the obtained values at different temperatures are collected in Table 2 and Fig. 7. Moreover, the overall formation constants β_i of the different intermetallic compounds have been summarized in Table 3.

The previous data can also be used to calculate the relative partial molar Gibbs free energies of Cd, $\Delta\tilde{G}_{\text{Cd}}$, and activities of Cd in the intermetallic compounds taken into account the reaction:



by means of the following equations³⁰

$$\Delta G_f^0(\text{HoCd}_x) = \Delta\tilde{G}_{\text{Ho}} + x\Delta\tilde{G}_{\text{Cd}} \quad (17)$$

$$\tilde{G}_{\text{Cd}} = RT \ln a_{\text{Cd}} \quad (18)$$

The obtained values are collected in Table 4.

Evaluation of Cd distillation at high temperature

Following the methodology proposed by Shibata *et al.*¹⁶ and Kurata *et al.*,³¹ based on the Scientific Group Thermodata Europe (SGTE) database,³² the Gibbs free energy of Ho–Cd intermetallic compounds, $G(\text{HoCd}_x)$, can be indicated with respect to the enthalpies of the pure elements in their reference states at 298.15 K, by the following equation:

$$\frac{\Delta G_f^0(\text{HoCd}_x)}{x+1} = \frac{G(\text{HoCd}_x)}{x+1} - \frac{1}{x+1} H_{\text{Ho}}^{\text{hcp},298\text{K}} - \frac{x}{x+1} H_{\text{Cd}}^{\text{liq},298\text{K}} \quad (19)$$

where $H_{\text{Ho}}^{\text{hcp},298.15\text{K}}$ and $H_{\text{Cd}}^{\text{liq},298.15\text{K}}$ are the enthalpies at 298.15 K in their reference states, -7.996 and -6.247 kJ mol⁻¹, respectively.³¹

Fig. 8 shows the obtained Gibbs free energy of Ho–Cd intermetallic compounds as a function of atomic fraction of Ho.

Although the decomposition temperature of HoCd₆ is not known, its Gibbs free energy has been represented assuming that HoCd₆ is not decomposed below 1000 K. Gibbs free energy

Table 3 Overall formation constants β_i of the different intermetallic

Intermetallic compound	Number of data	Formation reaction	β_i
HoCd ₆	98	$\text{Ho} + 6\text{Cd} \rightarrow \text{HoCd}_6$	$\beta_6 = \exp(-6.80 + 20\,329/T)$
HoCd _{45/11}	216	$\text{Ho} + \frac{45}{11}\text{Cd} \rightarrow \text{HoCd}_{\frac{45}{11}}$	$\beta_{45/11} = \exp(-5.54 + 18\,459/T)$
HoCd ₃	201	$\text{Ho} + 3\text{Cd} \rightarrow \text{HoCd}_3$	$\beta_3 = \exp(-4.73 + 16\,857/T)$
HoCd ₂	197	$\text{Ho} + 2\text{Cd} \rightarrow \text{HoCd}_2$	$\beta_2 = \exp(-3.65 + 14\,181/T)$
HoCd	135	$\text{Ho} + \text{Cd} \rightarrow \text{HoCd}$	$\beta_1 = \exp(-4.64 + 11\,911/T)$

Table 4 Thermodynamic properties of Cd for Ho–Cd intermetallic compounds in two-phase coexisting states at various temperatures

T/K	$\Delta\tilde{G}_{\text{Cd}}$ (kJ (mol Ho) ⁻¹)	a_{cd}
In the two-phases coexisting state between HoCd and HoCd₂		
673	-26.1 ± 1.1	(9.5 ± 1.8) × 10 ⁻³
698	-24.1 ± 0.7	(1.6 ± 0.2) × 10 ⁻²
723	-24.1 ± 0.8	(1.8 ± 0.2) × 10 ⁻²
748	-24.2 ± 1.3	(2.1 ± 0.4) × 10 ⁻²
773	-24.6 ± 1.3	(2.2 ± 0.4) × 10 ⁻²
823	-25.7 ± 0.7	(2.3 ± 0.2) × 10 ⁻²
In the two-phases coexisting state between HoCd₂ and HoCd₃		
673	-15.5 ± 1.4	(6.4 ± 1.5) × 10 ⁻²
698	-15.7 ± 0.5	(6.7 ± 0.6) × 10 ⁻²
723	-16.2 ± 0.8	(6.8 ± 0.9) × 10 ⁻²
748	-15.7 ± 0.8	(8.1 ± 1.1) × 10 ⁻²
773	-15.7 ± 0.9	(8.8 ± 1.0) × 10 ⁻²
823	-14.8 ± 0.9	(1.2 ± 0.1) × 10 ⁻¹
In the two-phases coexisting state between HoCd₃ and Ho₁₁Cd₄₅		
673	-9.1 ± 0.3	(2.0 ± 0.1) × 10 ⁻¹
698	-7.3 ± 0.4	(2.9 ± 0.2) × 10 ⁻¹
723	-8.0 ± 0.6	(2.7 ± 0.3) × 10 ⁻¹
748	-7.6 ± 0.6	(2.9 ± 0.3) × 10 ⁻¹
773	-7.4 ± 0.9	(3.2 ± 0.4) × 10 ⁻¹
823	-7.8 ± 0.6	(3.2 ± 0.3) × 10 ⁻¹
In the two-phases coexisting state between Ho₁₁Cd₄₅ and HoCd₆		
673	-4.2 ± 0.2	(4.7 ± 0.2) × 10 ⁻¹
698	-3.6 ± 0.5	(5.4 ± 0.5) × 10 ⁻¹
723	-4.1 ± 0.3	(5.0 ± 0.3) × 10 ⁻¹
748	-4.3 ± 0.4	(5.1 ± 0.3) × 10 ⁻¹
773	-4.1 ± 0.4	(5.4 ± 0.4) × 10 ⁻¹
823	-4.2 ± 0.5	(5.4 ± 0.4) × 10 ⁻¹

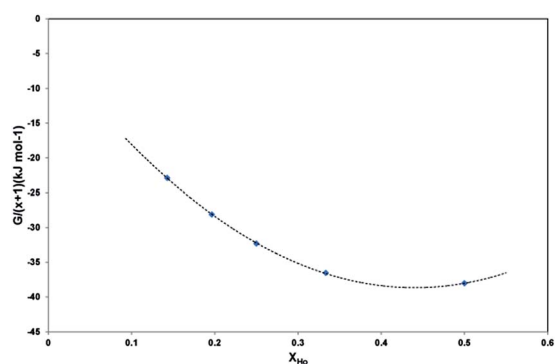


Fig. 8 Gibbs free energies of Ho–Cd intermetallics compounds at 973 K.

of Ho–Cd intermetallic compounds decreases gradually from HoCd₆ to HoCd and attains the minimum value at HoCd₂ and HoCd, suggesting that both intermetallics are relatively stable. Therefore, considering cathode product obtained by electro-refining of U–Pu–Zr metal fuel, significant amounts of U and Pu might be contained in the stables HoCd₂ and/or HoCd, which might affect their decomposition, a similar result to those reported by Shibata *et al.*¹⁶ in the stable CeCd₂.

Conclusions

It has been proved that the electro-reduction of Ho(III) at liquid Cd electrodes (*i.e.* Cd pool and Cd coated tungsten electrode) can be conveniently elucidated by electrochemical techniques. Due to a lowering of the activity of Ho in the metal phase, the redox potential of the Ho(III)/Ho couple at both electrodes was observed at more positive potential values than at inert ones.

CV, ASV, and OCP measurements using a Cd coated tungsten electrode have proven to be particularly useful for phase identification in intermetallic compounds. For avoiding the semi-infinite diffusion model, a CdFE has been chosen instead of a Cd pool, making possible the formation of distinct non-homogeneous phases during deposition.

Electromotive force, emf, measurements for various intermetallic compounds in two-phase coexisting states were carried out in the temperature range of 673–823 K. The activities and relative partial molar Gibbs energies of Ho were obtained from the measured emf for HoCd₆, HoCd_{45/11}, HoCd₃, HoCd₂ and HoCd intermetallic compounds. The formation energy of each intermetallic compound, and the global formation constants were also calculated. Moreover, the activities and relative partial molar Gibbs energies of Cd have also been calculated.

By extrapolating the molar Gibbs free energy of Ho–Cd intermetallic compounds to the Cd distillation temperature, it was observed that the molar Gibbs free energy of Ho in Ho–Cd intermetallic compounds decreases gradually from HoCd₆ to HoCd and attains to the minimum value at HoCd₂ and HoCd. This suggests on the Cd distillation from the U–Pu–Ho–Cd alloy that the dissolution of U or Pu into HoCd₂ and HoCd should be mostly taken into consideration.

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